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						Ì			Identifs	Chapi Multip
	9 Io A	8 B A H	7 B A &	6 В В	5 B A H	4 Fo	3 B A W	2. TI A B	Identify the choice that best completes the statement or answers the question 1 A measure of the ability of an atom in a chemical compound to A electron affinity. B electron configuration. D ionization	Chapter 12 Prep-Test Multiple Choice
	onic comp solid	low does It gai It giv It does Oxyg	What is the charge on the strontium ion? A 2- B 1-	low man	low man	or Groups plus 1. plus the	/hen chemic lost only.	he element oxygen, sodium.	measure electr	rep-To
	pounds	does oxygen obey the octet rule who It gains electrons. It gives up electrons. It does not change its number of elec Oxygen does not obey the octet rule.	ie charge	y valenc	y valenc	Groups 13 through 18, t plus 1. plus the period number.	n chemical co lost only. gained only.	nt that h en. m.	that best completes the teasure of the ability of electron affinity.	est
	are nom	obey thoons. cons. ange its	on the	e electro	e electro	ough 18, d numbe	mpound	nas the g	pletes the ability of ity.	
	ally in v	e octet n number y the oct	strontiur	ns does	ns are ir	the nun	s form,	reatest e	e statem f an aton	
	vhich ph	does oxygen obey the octet rule when realt gains electrons. It gives up electrons. It does not change its number of electrons. Oxygen does not obey the octet rule.	n ion?	How many valence electrons does a helium atom have? A 2 C 4 B 3 D 5	How many valence electrons are in an atom of phosphorus? A 2 C 4 B 3 D 5	ıber of v	valence	The element that has the greatest electronegativity is A oxygen. C o	ent or a	
<u> </u>	ysical st C D	reacting	D C	atom he C D	1 of phos C D	alence e C D	electrons C D	gativity i C D	nswers to emical co C D	
	ate at roc gas plasma	g to form	1+ 2+	ave?	sphorus? 4 5	lectrons is minus the minus 10.	s are those that shared only. lost, gained,	is chlorine. fluorine.	he questi ompound electron ionizati	
	Ionic compounds are normally in which physical state at room temperature? A solid C gas B liquid D plasma	How does oxygen obey the octet rule when reacting to form compounds? A It gains electrons. B It gives up electrons. C It does not change its number of electrons. D Oxygen does not obey the octet rule.				For Groups 13 through 18, the number of valence electrons is equal to the group number A plus I. C minus the period number, B plus the period number. D minus 10.	When chemical compounds form, valence electrons are those that may be A lost only. C shared only. B gained only. D lost, gained, or shared	in in	noice that best completes the statement or answers the question. A measure of the ability of an atom in a chemical compound to attract electrons is called A electron affinity. C electronegativity. B electron configuration. D ionization potential.	
	erature?	ınds?				to the gr d numbe	ay be shared.		ct electro y. ıtial.	
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		Which m A SO ₂ B CO ₃	Use V: A tei B lin	VSEPI A th B th	What is A A B B	A. Cl—H:	A pol B ion	A ele B pr	An Ion A ac B va	Which A pro
1)		ı molecu D ₂	VSEPR the tetrahedral linear	VSEPR theory is a model for predicting A the strength of metallic bonds. B the shape of molecules.	s the Lev	—H:	polar covalent. ionic.	electrons. protons.	An Ionic bond is a bond between. A a cation and an anion B valence electrons and cations A chemical bond results from the	Which of the following particles are free to drift in metals? A protons C neutron B electrons D cations
		molecule is linear? h h	ory to p	is a mod h of met of molec	vis struc	B . :H—	ent.		s a bond d an anio ctrons a	llowing
		ear?	redict th	el for pr allic bon ules.	the Lewis structure for hydrogen chloride, HCl? C C D D	- Q:	on open of		betweer on d catior	particles
1)		e molecı	edicting ds.	nydroger	C. :H—	the bond		ı .	are free
2	-	DC	ılar shape C D	DC	ı chloride C D	т С С С	C C	חם	attractio	to drift i C D
7		ChO	e of the bent trigo		e, HCl?	нсі 0 ₂ р .	nonp pure	neutrons. dipoles	the ic	in metals? neutrons cations
- Company		Ü	f the carbon tet bent trigonal planar	lattice energy values.		н—сі:	nonpolar covalent. pure covalent. he atoms NOT a no	ons.	ns of tw	ons
			etraiodid	values.		-	elent.		An Ionic bond is a bond between A a cation and an anion C the ions of two different I B valence electrons and cations D the ions of two different I A chemical bond results from the mutual attraction of the nuclei of atoms and	
L			Use VSEPR theory to predict the molecular shape of the carbon tetraiodide molecule, CL, A tetrahedral C bent B linear D trigonal planar				ar covalent. D pure covalent. D pure covalent. D pure covalent. D pure covalent.	F .	the ions of two different metals the ions of two different nonmetals of the nuclei of atoms and	
1			ıle, CL.				A polar covalent. A polar covalent. D pure covalent. D pure covalent. D pure covalent. D pure covalent. A polar covalent. D pure covalent. A polar covalent be a point covalent. A polar covalent bond between the atoms NOT a nonpolar covalent bond.);·	etals	

- | 19| What is the molecular geometry of this lewis structure (NF₃)?

 A square pyramidal C trigonal pyramidal B tetrahedral D linear C trigonal pyramidal D linear
- Which of the following compounds has an ionic bond? C NaCl D CO2
- 21 Which of the following is the best representation of the compound rubidium chloride?

- 22 How many valence electrons are in a Se² ion?

 A 8
 C
 B 7
 D
- electronegativities.
- B Polar because H and F are both nonmetals.
- D
- In the Molecular Models Lab, you built the structure for Hydrogen Fluoride. What is the polarity of that molecule and why?
- A Nonpolar because H and F have the same C Polar because F is more electronegative
- Nonpolar because H is more electronegative than F. than H.

Matching

Match each item with the correct statement below.

A halide ion
E
B octet rule
C ionic bond
G
D Lewis dot structure
H

valence electron acid base reaction metallic bond bent molecule

24 an electron in the highest occupied energy level of an atom

25 Atoms react so as to acquire the stable electron structure of a noble gas.

26 a depiction of valence electrons around the symbol of an element

27 the attraction of valence electrons for metal ions

28 the force of attraction binding oppositely charged ions together

29 Group 7 (or 17) anions

30 H₂O