

**Nomenclature Type I and II --No Molecules**

NAME \_\_\_\_\_  
Period \_\_\_\_\_  
Date \_\_\_\_\_

**GIVE THE CHEMICAL FORMULA FOR THE FOLLOWING**

1. Sodium Bromide 1. \_\_\_\_\_
2. Iron (III) Chloride 2. \_\_\_\_\_
3. Zinc Oxide 3. \_\_\_\_\_
4. Calcium Sulfide 4. \_\_\_\_\_
5. Zinc Phosphate 5. \_\_\_\_\_
6. Ammonium Sulfide 6. \_\_\_\_\_
7. Magnesium bicarbonate 7. \_\_\_\_\_
8. Copper (III) oxide 8. \_\_\_\_\_
9. Sodium nitrate 9. \_\_\_\_\_
10. Lead (II) nitrate 10. \_\_\_\_\_
11. lithium chloride 11. \_\_\_\_\_
12. Iron (II) phosphate 12. \_\_\_\_\_

**NAME THE FOLLOWING COMPOUNDS AND MOLECULES**

13. MgS 13. \_\_\_\_\_
14. FeS 14. \_\_\_\_\_
15. Al<sub>2</sub>O<sub>3</sub> 15. \_\_\_\_\_
16. (NH<sub>4</sub>)<sub>3</sub> PO<sub>4</sub> 16. \_\_\_\_\_
17. KNO<sub>3</sub> 17. \_\_\_\_\_
18. Na<sub>2</sub>CrO<sub>4</sub> 18. \_\_\_\_\_
19. CaO 19. \_\_\_\_\_
20. CuCl 20. \_\_\_\_\_
21. Na<sub>2</sub>SO<sub>3</sub> 21. \_\_\_\_\_
22. MgSO<sub>4</sub> 22. \_\_\_\_\_
23. H<sub>2</sub>SO<sub>3</sub> 23. \_\_\_\_\_
24. H<sub>2</sub>S 24. \_\_\_\_\_
25. KMnO<sub>4</sub> 25. \_\_\_\_\_
26. CaCl<sub>2</sub> 26. \_\_\_\_\_
27. H<sub>2</sub>SO<sub>4</sub> 27. \_\_\_\_\_
28. CaSO<sub>3</sub> 28. \_\_\_\_\_
29. How many electrons and protons does the fluoride ion have? 29. e<sup>-</sup> \_\_\_\_\_  
p<sup>+</sup> \_\_\_\_\_