IB/AP Chemistry Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Period \_\_\_ Date \_\_\_/\_\_\_/\_\_\_

1 • Matter and Measurement- BIG IDEA 2

**CHROMATOGRAPHY LABETTE**

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| **To Do:**  1. Set up apparatus as shown:   * glass stirring rod * piece of tape * strip of chromatography paper * scissors * cup |  |
| 2. Draw a ***pencil*** line near the bottom of the chromatography paper. Why use a pencil? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  3. Dot samples from two different black felt pens on the pencil line.  4. Add ***just enough water*** so the tip of the paper is wet.  **To Notice:**  1. What happens to the ink spots when the water moves through them?  2. Is the ink a pure substance or a mixture?  3. Look at the two ink spots and those of your neighbors. Are all inks the same mixture? \_\_\_\_ Justify your answer:  4. A big idea in this chapter is that mixtures can be separated by exploiting differences in physical properties. The two properties in this case are called ***solubility*** and ***adsorption***.  Define these terms:   * solubility * adsorption | |

5. How much would the component spot travel if the component is very soluble.



6. How much would the component spot travel if the component strongly adsorbs.

