

Acids and Bases Review Worksheet

1. Why are acids donors
2. In a base the number of H_3O^+ ions is _____ the number of OH^- ions.
3. In water, how does the concentration [] of H_3O^+ compare to the amount of OH^- ?
4. We discussed how acids and bases are able to conduct electricity. Why is this?
5. What does pH literally measure?
6. Draw a pH scale. Be sure to include the numeric range as well as indicate which area is acidic, basic, and neutral.
7. What does it mean, in terms of ion concentration, if a solution is 'neutral'?
8. In our antacid lab, what product did we use to create an acid?
9. What was the overall purpose of our antacid lab? Be specific.
10. As the concentration of H^+ (or H_3O^+) increases in a solution, what happens to the pH?
11. Identify each of the following as an Acid or a Base?
 - Lemon juice is _____
 - Detergent is _____
 - Bleach is _____
 - Aspirin is _____
12. Place the following in order from strongest base to strongest acid?
 - Tomatoes - Sea water
 - Lime - Clean Rain
13. What are the products when an acid and base are added together?
14. What are indicators?
15. What is an aqueous solution?
16. What does it mean when a solution is neutral?
17. What is the difference between a strong acid and a weak acid?
18. What is the difference between a strong base and a weak base?
19. What is an electrolyte
20. What is ionization?
21. What is neutralization?
22. Name some acids and bases found in the home.
23. What are the properties of acids?
24. What are the properties of bases?
25. The label on a bottle indicates that the substance inside has a **pH of 14**. This tells you that the substance is a _____
26. Baking soda has a pH of 9, and household ammonia has a pH of 12. This means that ammonia is _____ times more basic than baking soda.

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