Polyatomic ions

OTHER CATIONS				**OTHER ANIONS**		Cyanide	CN^-
Ammonium	NH_4^+			Acetate	$C_2H_3O_2^-$	Hydroxide	OH^-
Hydronium	H_3O^+			Chromate	CrO_4^{2-}	Oxalate	$C_2O_4^{2-}$
				Dichromate	$Cr_2O_7^{2-}$	Peroxide	O_2^{2-}
Perbromate	BrO_4^-	Perchlorate	ClO_4^-	Periodite	IO_4^-		
Bromate	BrO_3^-	Chlorate	ClO_3^-	Iodate	IO_3^-	Nitrate	NO_3^-
Bromite	BrO_2^-	Chlorite	ClO_2^-	Iodite	IO_2^-	Nitrite	NO_2^-
Hypobromite	BrO^-	Hypochlorite	ClO^-	Hypoiodite	IO^-		
Carbonate	CO_3^{2-}	Sulfate	SO_4^{2-}	Phosphate	PO_4^{3-}	Permanganate	MnO_4^-
		Sulfite	SO_3^{2-}	Phosphite	PO_3^{3-}		
BI-ANIONS				**THIO-ANIONS**			
Bicarbonate	HCO_3^-			Thiocyanate	SCN^-		
Bisulfate	HSO_4^-			Thiosulfate	$S_2O_3^{2-}$		
Bisulfide	HS^-						
Bisulfite	HSO_3^-						

Names of Some Common Acids	
HNO_3	Nitric acid
HNO_2	Nitrous acid
H_2SO_4	Sulfuric acid
H_2SO_3	Sulfurous acid
H_3PO_4	Phosphoric acid
H_2CO_3	Carbonic acid
CH_3COOH	Acetic acid
HCl	Hydrochloric acid
HF	Hydrofluoric acid
HBr	Hydrobromic acid

Mechanisms and rules for writing chemical formulas:

THE CRISS-CROSS METHOD

RULE 1: The resulting formula for a compound must have a total charge of zero (0).
RULE 2: Write the positive ion first and cross the valences.
RULE 3: Do not cross any signs.
RULE 4: Don't cross any ones.
RULE 5: If both valences are the same, don't cross them.
RULE 6: More than one atom, more than one time, use parentheses
RULE 7: If the final answer has subscripts that can be reduced, they must be reduced.
RULE 8: If the name of the compound has prefixes in it, change the prefixes to subscripts and do not cross the valences.