		Increasing electronegativity																
						H 2.1												
	vity	Li 1.0	Be 1.5			2.1								B 2.0	C 2.5	N 3.0	O 3.5	F 4.0
	Decreasing electronegativity	Na 0.9	Mg 1.2											Al 1.5	Si 1.8	P 2.1	S 2.5	Cl 3.0
	ctron	K 0.8	Ca 1.0	Sc 1.3	Ti 1.5	V 1.6	Cr 1.6	Mn 1.5	Fe 1.8	Co 1.9	Ni 1.9	Cu 1.9	Zn 1.6	Ga 1.6	Ge 1.8	As 2.0	Se 2.4	Br 2.8
	ng elc	Rb 0.8	Sr 1.0	Y 1.2	Zr 1.4	Nb 1.6	Mo 1.8	Tc 1.9	Ru 2.2	Rh 2.2	Pd 2.2	Ag 1.9	Cd 1.7	In 1.7	Sn 1.8	Sb 1.9	Te 2.1	I 2.5
	reasi	Cs 0.7	Ва 0.9	La-Lu 1.0-1.2	Hf 1.3	Та 1.5	W 1.7	Re 1.9	Os 2.2	Ir 2.2	Pt 2.2	Au 2.4	Hg 1.9	Tl 1.8	Pb 1.9	Bi 1.9	Po 2.0	At 2.2
	Dec	Fr 0.7	Ra 0.9	Ac 1.1	Th 1.3	Pa 1.4	U 1.4	Np-No 1.4-1.3										

Key

< 1.5

1.5-1.9

2.0 - 2.9

3.0-4.0

If the <u>difference</u> in the Electronegativities is:
• Equal to **0** it is a <u>Covalent Bond</u>

- Greater than **0** but Less than **2** it is a **Polar Covalent Bond**
- Greater than **2** it is an **lonic Bond**