## **Nuclear Balancing Act**

| Name_  | <br> |
|--------|------|
| Period | <br> |
| Date   |      |

Write a balanced equation for each of the following nuclear reactions

Example: Chlorine-36 decays by beta emission

 $^{36}_{17}\text{Cl} \rightarrow ^{36}_{18}\text{Ar} + ^{0}_{-1}\text{e}$ 

- 1. Krypton-87 decays by beta emission.
- 2. Curium-240 decays by alpha emission.
- 3. Uranium-232 decays by beta decay.
- 4. Silicon-32 decays by beta emission.
- 5. Americium-243 decays by alpha emission.
- 6. Boron-8 decays by positron emission.
- 7. Iridium-192 decays by positron emission.
- 8. Germanium-68 undergoes electron capture.
- 9. Fluorine-18 undergoes electron capture.
- 10. Toughie! Lead-210 decays by emitting both a beta and an alpha particle.