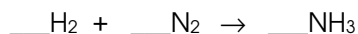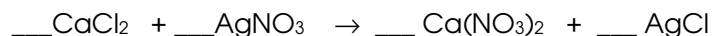# Mass Practice Worksheet

$$\_\_\_H_2 + \_\_\_N_2 \rightarrow \_\_\_NH_3$$

1. How many moles of H_2 are needed to react with 2.5 moles of N_2?

2. How many moles of NH_3 can be produced by 2.5 moles of N_2?

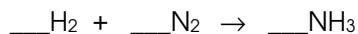$$\_\_\_CaCl_2 + \_\_\_AgNO_3 \rightarrow \_\_\_ Ca(NO_3)_2 + \_\_\_ AgCl$$

3. What mass of $AgNO_3$ solution is needed to react with 2.50 g or $CaCl_2$?

4. What is the mass of $CaCl_2$ needed to react with 0.250 g $AgNO_3$?

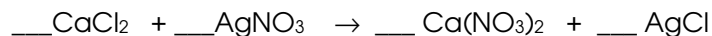$$\_\_\_HCl + \_\_\_ NaHCO_3 \rightarrow \_\_\_NaCl + \_\_\_ H_2O + \_\_\_ CO_2$$

5. What mass of NaCl can be produced from 2.00 g of $NaHCO_3$ in the above reaction?

6. What mass of CO_2 can be produced by reacting 2.50 grams of $NaHCO_3$ with excess acid?

Mass Practice Worksheet

___H_2 + ___N_2 → ___NH_3

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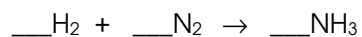___$CaCl_2$ + ___$AgNO_3$ → ___ $Ca(NO_3)_2$ + ___ $AgCl$

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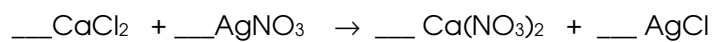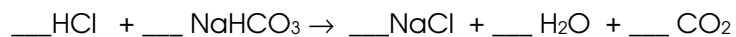___HCl + ___ $NaHCO_3$ → ___$NaCl$ + ___ H_2O + ___ CO_2

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