Empirical Formula WS

1. Write the Empirical Formula for Each of the Following:

    a. P_4O_6 c. CH_2OHCH_2OH e. $C_6H_8O_6$ g. $Cu_2C_2O_4$

    b. C_6H_9 d. $BrCl_2$ f. $C_{10}H_{22}$ h. Hg_2F_2
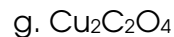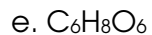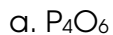
2. Write the Empirical Formula for Each of the Following:

    a. A compound composed of: 72% iron and 27.6% oxygen by mass.

    b. A compound composed of: 9.93% carbon, 58.6% chlorine, and 31.4% fluorine.

    c. A compound composed of: .556g carbon and .0933g hydrogen.

3. Write the Molecular Formula for Each of the Following:

    a. A compound with a molecular mass of 70g/mol and an empirical formula of CH_2.

    b. A compound with a molecular mass of 46.0g/mol and an empirical formula of NO_2