

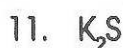
NAMING BINARY COMPOUNDS (IONIC)

Name _____

Name the following ionic compounds using Roman numerals where necessary.



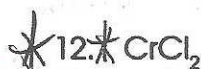
Barium Chloride



Potassium Sulfide



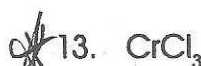
Sodium Fluoride



Chromium (II) Chloride



Silver Oxide



Chromium (III) Chloride



Copper (I) Bromide



Calcium Oxide



Copper (II) Bromide



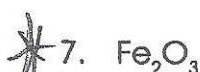
Barium Phosphide



Iron (II) Oxide



Mercury (I) Iodide



Iron (III) oxide



Sodium Oxide



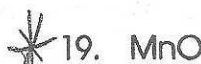
Magnesium Sulfide



Beryllium Sulfide



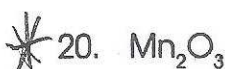
Aluminum Oxide



Manganese Oxide



Calcium Iodide



Manganese Oxide

NAMING IONIC COMPOUNDS

Name KEY

Name the following compounds using the Stock Naming System.

- | | |
|--|----------------------------------|
| 1. CaCO_3 | <u>Calcium Carbonate</u> |
| 2. KCl | <u>Potassium Chloride</u> |
| * 3. FeSO_4 | <u>Iron (II) Sulfate</u> |
| 4. LiBr | <u>Lithium Bromide</u> |
| 5. MgCl_2 | <u>Magnesium Chloride</u> |
| * 6. FeCl_3 | <u>Iron (III) Chloride</u> |
| 7. $\text{Zn}_3(\text{PO}_4)_2$ | <u>Zinc Phosphate</u> |
| 8. NH_4NO_3 | <u>Ammonium Nitrate</u> |
| 9. $\text{Al}(\text{OH})_3$ | <u>Aluminum Hydroxide</u> |
| * 10. $\text{CuC}_2\text{H}_3\text{O}_2$ | <u>Copper (II) Acetate</u> |
| * 11. PbSO_3 | <u>Lead (II) Sulfite</u> |
| 12. NaClO_3 | <u>Sodium Chlorate</u> |
| 13. CaC_2O_4 | <u>Calcium Oxalate</u> |
| * 14. Fe_2O_3 | <u>Iron (III) Oxide</u> |
| 15. $(\text{NH}_4)_3\text{PO}_4$ | <u>Ammonium Phosphate</u> |
| 16. NaHSO_4 | <u>Sodium Bisulfate</u> |
| * 17. Hg_2Cl_2 | <u>Mercury (I) Chloride</u> |
| 18. $\text{Mg}(\text{NO}_2)_2$ | <u>Magnesium Nitrite</u> |
| * 19. CuSO_4 | <u>Copper (II) Sulfate</u> |
| 20. NaHCO_3 | <u>Sodium Hydrogen Carbonate</u> |
| * 21. NiBr_3 | <u>Nickel (III) Bromide</u> |
| 22. $\text{Be}(\text{NO}_3)_2$ | <u>Beryllium Nitrate</u> |
| 23. ZnSO_4 | <u>Zinc Sulfate</u> |
| * 24. AuCl_3 | <u>Gold (III) Chloride</u> |
| 25. KMnO_4 | <u>Potassium Permanganate</u> |